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Chapter 12 REVIEW:
Stoichiometry,
Theoretical, Actual &
Percent yield Part I.
Stoichiometry 1. 1 ZnI₂

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→ 1 Zn + 1 I₂ How many grams of iodine will you produce if you begin your rxn with 56.7g of zinc iodide?
56.7 g 1 mole 1 mole
253.8 g 319.19 g 1 mole 1 mole = 45.08 grams of Iodine 2. 1
 $\text{Pb}_2(\text{CO}_3)_3 + 6 \text{K} \rightarrow 2 \text{Pb} + 3 \text{K}_2\text{CO}_3$

Chapter 12 REVIEW: Stoichiometry, Theoretical, Actual

...

Chapter 12 Review
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Stoichiometry

Important Vocabulary •

Stoichiometry- The calculations of quantities in chemical reactions in a subject of chemistry • Mole ratio- a conversion factor derived from the coefficients of a balanced chemical equation interpreted in terms of moles

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Chapter 18

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Key Chapter 12
Stoichiometry Section
Review In Example
12.2.1 and Example
12.2.2, the identity of
the limiting reactant
has been apparent:
[Au(CN) 2] -, LaCl 3,
ethanol, and para-

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nitrophenol. When the limiting reactant is not apparent, we can determine which

Chapter 12 Stoichiometry Section Review Answer Key

Chapter 12 Review.
Stoichiometry.

Important Vocabulary.

- Stoichiometry- The calculations of quantities in chemical reactions in a subject of chemistry
- Mole

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ratio- a conversion factor derived from the coefficients of a balanced chemical equation interpreted in terms of moles • Limiting reagent- The reactant that determines the amount of product that can be formed by a reaction • Excess reagent- The reaction that is not completely used up in a reaction • Theoretical yield- Maximum ...

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Ms_Li_ClearView.

Chemistry: Chapter 12:
Stoichiometry.

Stoichiometry. Is the
number of atoms
conserved in a b.... Is
the number of
molecules conserved
in.... Is the number of
moles conserved in a
b.... The calculation of
quantities in chemical
reactions. Yes... $N_2 +$
 $3H_2 \rightarrow 2NH_3$... 2

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atoms + 6 atoms \rightarrow 8
atoms.

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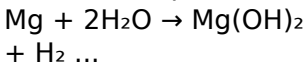
Overview of Chemistry
1 Honors Chapter 12:
Stoichiometry. Terms
in this set (21)

Stoichiometry. The
calculation of
quantities in chemical
reactions is a subject of
chemistry. Mole ratio.

... Use the following

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balanced equation to
answer the question:



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Stoichiometry 379

CHAPTER 12

Assessment 36. a. Two formula units KClO_3

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decom-
pose to form
two formula units KCl
and three molecules
O₂. b. Four molecules
NH₃ react with six
molecules NO to form
five mol-ecules N₂ and
six molecules H₂O. c.
Four atoms K react
with one molecule O₂
to form two formula
units K₂O. 37. a. Two
mol KClO₃ decompose
to

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Chapter 12
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Stoichiometry 12.1 The
Arithmetic of Equations

12.2 Chemical
Calculations 12.3

Limiting Reagent and
Percent Yield . . .

excess of 1800 is a
logical answer. • The
unit of the known (FSW
3 HP 2) cancels. • The
answer has the correct
unit (W). 3 Evaluate
Does the result make
sense?

Chapter 12

Prentice Hall Chemistry

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Chapter 12:
Stoichiometry Chapter
Exam Take this
practice test to check
your existing
knowledge of the
course material. We'll
review your answers
and create a Test Prep
...

Prentice Hall Chemistry Chapter 12: Stoichiometry ...

Chapter 12 Multiple
Choice Identify the
letter of the choice that

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best completes the statement or answers the question. ____ 1. The first step in most stoichiometry problems is to ____.

- a. add the coefficients of the reagents
- c. convert given quantities to volumes
- b. convert given quantities to moles
- d. convert given quantities to masses

____ 2.

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Prioritize Schedule your
time realisti-cally

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key. Stick to your
deadlines. . . . 378

Chapter 12 12CHAPTER

Study Guide Key

Concepts 12. 1 The

Arithmetic of Equations

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